**Procedure for Quenching Organic Peroxides and Hydroperoxides**

Reference 1: Peroxides and peroxide-forming compounds  
Clark, Donald E Chemical Health and Safety, September-October 2001, Vol.8(5), pp.12-22. [https://doi.org/10.1016/S1074-9098(01)00247-7](https://doi.org/10.1016/S1074-9098%2801%2900247-7)

Spilled peroxides should be absorbed on vermiculate as soon as possible. If appropriate facilities are available, the vermiculate-peroxide mixture can be incinerated directly or may be slurried by stirring with a suitable solvent. The slurry can be treated with an acidic ferrous sulfate solution (60 g ferrous sulfate + 6 mL con sulfuric acid + 110 mL water). Never flush organic peroxides down the drain.

Reference 2: [Prudent Practices in the Lab: Handling and Disposal of Chemicals](http://www.nap.edu/catalog.php?record_id=4911)

The section on peroxides is Chapter 7.D.2.5 pg 162-163 (<http://www.nap.edu/openbook.php?record_id=4911&page=162#p200063c99970162003>)

"Removal of peroxides with ferrous sulfate:

http://www.nap.edu/books/0309052297/xhtml/images/img00014.gif

A solution of 6 g of FeSO4 · 7H2O, 0.6 mL of concentrated sulfuric acid, and 11 mL of water is stirred with 1 L of water-insoluble solvent until the solvent no longer gives a positive test for peroxides. Usually only a few minutes are required.

Diacyl peroxides can be destroyed by this reagent as well as by aqueous sodium hydrogen sulfite, sodium hydroxide, or ammonia. However, diacyl peroxides with low solubility in water, such as dibenzoyl peroxide, react very slowly. A better reagent is a solution of sodium iodide or potassium iodide in glacial acetic acid.

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