**Procedure for Quenching Organic Peroxides and Hydroperoxides**

(Reference: [Prudent Practices in the Lab: Handling and Disposal of Chemicals](http://www.nap.edu/catalog.php?record_id=4911))

"Removal of peroxides with ferrous sulfate:

http://www.nap.edu/books/0309052297/xhtml/images/img00014.gif

A solution of 6 g of FeSO4 · 7H2O, 6 mL of concentrated sulfuric acid, and 11 mL of water is stirred with 1 L of water-insoluble solvent until the solvent no longer gives a positive test for peroxides. Usually only a few minutes are required.

Diacyl peroxides can be destroyed by this reagent as well as by aqueous sodium hydrogen sulfite, sodium hydroxide, or ammonia. However, diacyl peroxides with low solubility in water, such as dibenzoyl peroxide, react very slowly. A better reagent is a solution of sodium iodide or potassium iodide in glacial acetic acid.

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The section on peroxides is Chapter 7.D.2.5 pg 162-163 (<http://www.nap.edu/openbook.php?record_id=4911&page=162#p200063c99970162003>)

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